

Water Chemistry of the St. Lawrence River

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With a mean discharge of $14000 \text{ m}^3/\text{sec}$, the St. Lawrence River ranks eleventh among world rivers (Inman and Nordstrom, 1971). Much of the precipitation in the drainage basin occurs in the form of snow; and lakes covering glacial topography make up to 29% of the total drainage area of 50000 km^2 (d'Anglejan and Smith, 1973; Loring and Nota, 1973). Since the river is flowing out of the Great Lakes, the geology of the water shed is complicated. All the three major rock types are represented in the water shed as well as in the drainage basin. The Canadian Shield and the Appalachian Mountains contribute to the hard-rock part of the basin while the glacial sediments from the Wisconsin age and the Paleozoic Strata constitute the soft-rock part.

In addition to the Great Lakes, the major source of water is the tributaries such as the Ottawa, Richelieu and Saguenay Rivers. As shown by Gibbs (1967) for the Amazon river, the geological complexity of the drainage basin is likely to be reflected in the water chemistry of the St. Lawrence River. Except for the water analyses compiled by Livingstone (1960), no systematic study has so far been made for the St. Lawrence River.

In this report, water quality data obtained from the National Water Quality Data Bank (NAQUADAT) are analyzed graphically to show some of the inter-relationship among different variables. Equilibrium composition of the suspended matter has been derived theoretically from the water analyses, and possible water-sediment reactions are speculated.

RESULTSStation Location:

Fig. 1 shows the location of the stations selected for the purpose of the present work. Among a large number of stations for which water quality data are available, only a few are chosen so as to be representative of the entire length of the river. Work is currently under progress in the estuary region of the river and the present report is addressed only to the fresh-water region of the St. Lawrence River.

Major Constituents:

Fig. 2 shows the values of total dissolved load at various points between Montreal and Quebec City. The vertical bars in the figure represent seasonal variations. Because of discharge effects, total dissolved load can be expected to vary seasonally (Edwards, 1973; Gibbs, 1967, 1972; Grove, 1972; Livingstone, 1960). Longitudinal variations in the dissolved load can be

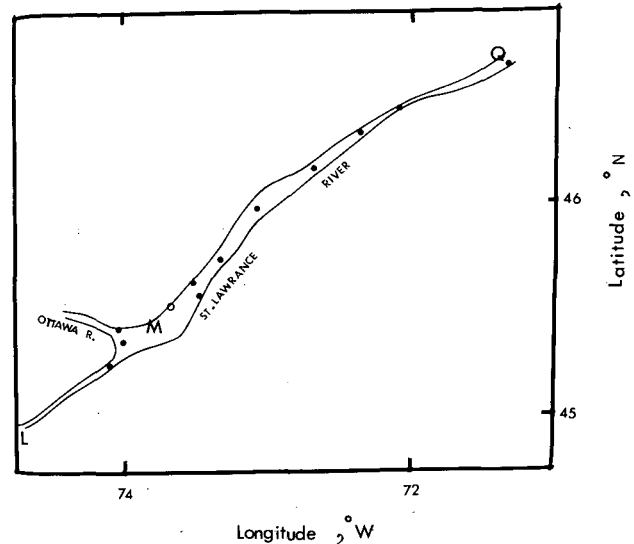


FIG. 1: Location of stations along the St. Lawrence River. M-Montreal; Q-Quebec City; L-Lake Ontario.

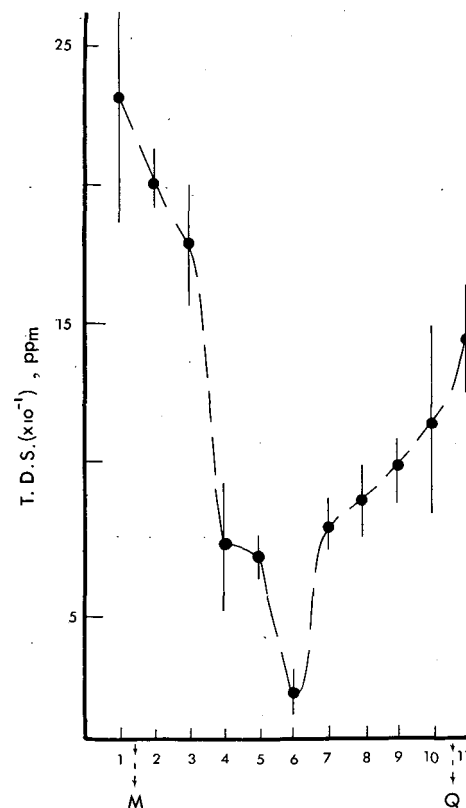


FIG. 2: Seasonal and longitudinal variations in the total dissolved solids (T.D.S.). Dots represent arithmetic mean.

explained by either dilution at localities where a tributary with a small dissolved load may enter the main river, or by removal of some salts by the ecosystem or by inorganic precipitation. On the other hand, increase in the dissolved load can be explained by the addition of salts through wastes from industries located along the length of the river, or by dissolution of some of the inorganic phases in the sediments, or by addition of water from a tributary draining a soft rock basin. In the case of the St. Lawrence river, the longitudinal distribution passes through a minima at a point roughly half way between Montreal and Quebec City. The removal of more than 75 per cent of the dissolved load at this point is due probably to the entrapment of the river water in a lake-like system present there, or due to the sample being taken at the junction of two small tributaries draining the Canadian Shield area. Heavy industrialization and urbanization all along the length of the Ottawa river, around Montreal and also around Quebec city may account for the higher total dissolved load observed in these regions.

The world average discharge of dissolved load is only 117 ppm (Gibbs, 1972) or 120 ppm (Livingstone, 1960). Even the tropical African rivers such as the Niger and Senegal Rivers have a maximum dissolved load of only 80 ppm (Grove, 1969); the Ganges and Brahmaputra Rivers, before reaching their subtropical and sedimentary basin, carry a maximum of only 140 ppm and 132 ppm respectively (Government of India Report, 1955; Handa, 1972). The Amazon River, whose dissolved load is derived from mountain areas in the drainage basin, carries only 62 ppm (Gibbs, 1967, 1972). Thus the mass transfer of the total dissolved load from the water shed and the drainage basin by the St. Lawrence River to the Atlantic Ocean may be very significant, and may be important in the marine chemistry of the North Atlantic Ocean.

Figure 3 shows the variations in the concentrations of major ions with the total dissolved load. From the fresh water towards the estuary, the predominating ions in solution will change from Ca^{++} and HCO_3^- to Na^+ and Cl^- . Since all the stations considered in this report are located within the fresh water region of the St. Lawrence river, changes in the ionic ratios cannot be explained by sea-water mixing. In Figure 3 it is seen that the ratio $\text{Na}/\text{Na}+\text{Ca}$ sharply increases with the total dissolved salts around Montreal whose position is indicated with an arrow in the figure; below the arrow, stations are located between Lake Ontario and Montreal and above that point, the stations are located between Montreal and Quebec City. The sharp increases of Na^{++} accompanied by a corresponding decrease in Ca^{++} is probably due to winter salts washed off from the urban areas into the river. Ruthford (1972) has observed a high percentage of Na^+ in the sediment waters from the bottom sediments of St. Lawrence river near Kingston, Ontario. The behaviour of the anionic ratios, $(\text{Ca HCO}_3/\text{Na}+\text{Cl})$, shown in the same figure, suggest that any increase in Na^+ is also accompanied by increase in Cl^- ; NaCl used during the winter months on the streets offers a convenient source for both Na^+ and Cl^- . The Champlain Sea sed-

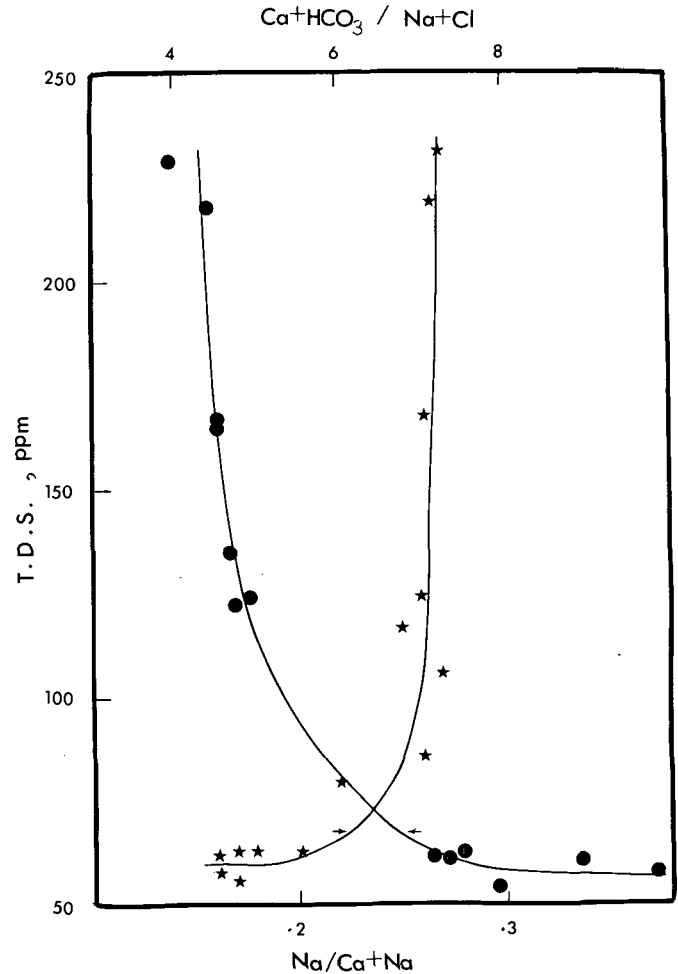


FIG. 3: Variations in major ionic ratios with the total dissolved solids. Stars = $\text{Ca}+\text{HCO}_3/\text{Na}+\text{Cl}$; Dots = $\text{Na}/\text{Na}+\text{Ca}$. Arrow indicates location of Montreal.

iments, part of the drainage basin of the St. Lawrence River, could be a source for many ions in the water, particularly for stations between Montreal and Quebec city.

Figure 4 shows the relationship between the total dissolved solids and the major anions in the river water. The behaviour of Cl/HCO_3 and Cl can be predicted from the previous figure since Na^+ and Cl^- , on the one hand, and Ca^{++} and HCO_3^- on the other, could be derived from common sources. Continuous increase in the SO_4^{--} concentrations with the total dissolved salts suggests, but does not prove, that they have common controlling factors. Sharp increase in SO_4^{--} near Montreal may be due to the industrial effluents from refineries. Pulp mills, petrochemicals and chemical industries are potential contributors of sulphur-bearing materials to the river waters (Gove and Gellman, 1972; Mytelka, 1972).

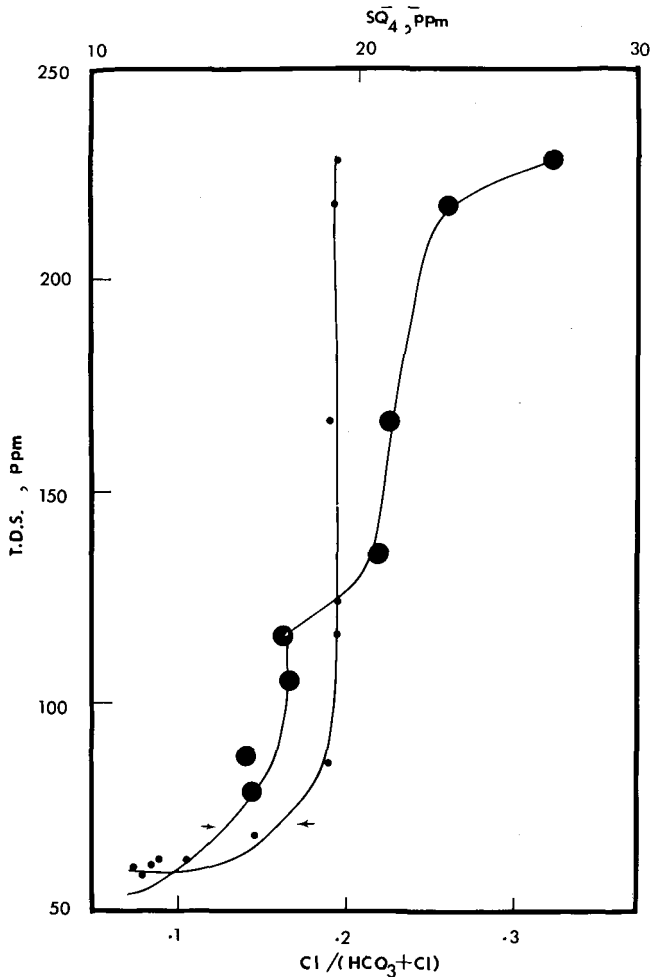


FIG. 4: Variations in anionic ratios with total dissolved solids (T.D.S.). Small Dots= $Cl-HCO_3+Cl$; Large Dots= SO_4 . Arrow indicates location of Montreal.

Figure 5 shows the relationship between total reactive silica and pH ranges, reactive silica can be considered to represent all the dissolved silica species in fresh waters. In silicate reactions, pH is an important variable (Garrels and Christ, 1965); based on solubility reasoning, acidic waters should be silica deficient. Figure 5 indicates the role of two factors in regulating the silica levels in the river water: (a) Seasonal variation - The silica numbers have been grouped into seasonal ranges and the highest values are observed during the colder months of February-March; lowest values are observed in the Spring and Summer times of May-October. (b) pH variation - While a definite and clearcut relationship is lacking between silica and pH, a negative correlation is suggested by the plot obtained from the analytical data.

Silicate minerals are more soluble at higher pH (Mackenzie and Garrels, 1965). If the silica content in the St. Lawrence river is controlled primarily by silicate reactions, then the plot should have been reversed; it is hence likely that factors other than mineral-water reactions are responsible for the observed increase in silica at lower pH values; the low silica values at higher pH may be due to sorption of silica by clay minerals, as was shown by Siever (1973) in laboratory experiments. Organic productivity is also an additional variable in silica chemistry: the small amount of silica at higher pH values probably reflects the uptake of silica by organisms in the warm months of June to October. It is also possible that the higher silica values reported at lower pH may represent analysis of colloidal silica (Siever, 1962; McKeague and Cline, 1963). Even at the highest silica concentrations, the St. Lawrence River water is under-saturated with respect to both quartz and amorphous silica (Helgeson, 1969) and

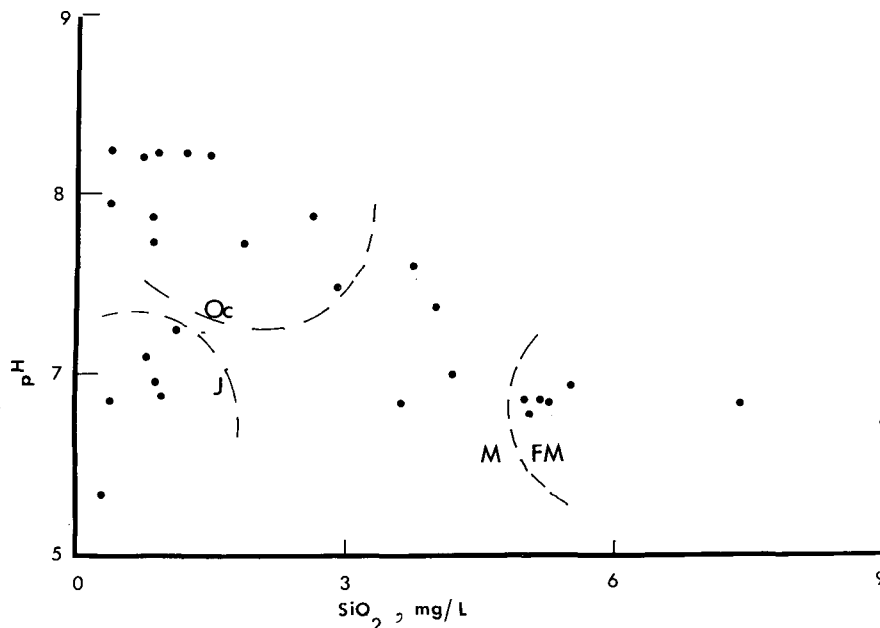


FIG. 5: Variations in Silica contents with pH. Oc-October; J-June; M-May; FM-February-March.

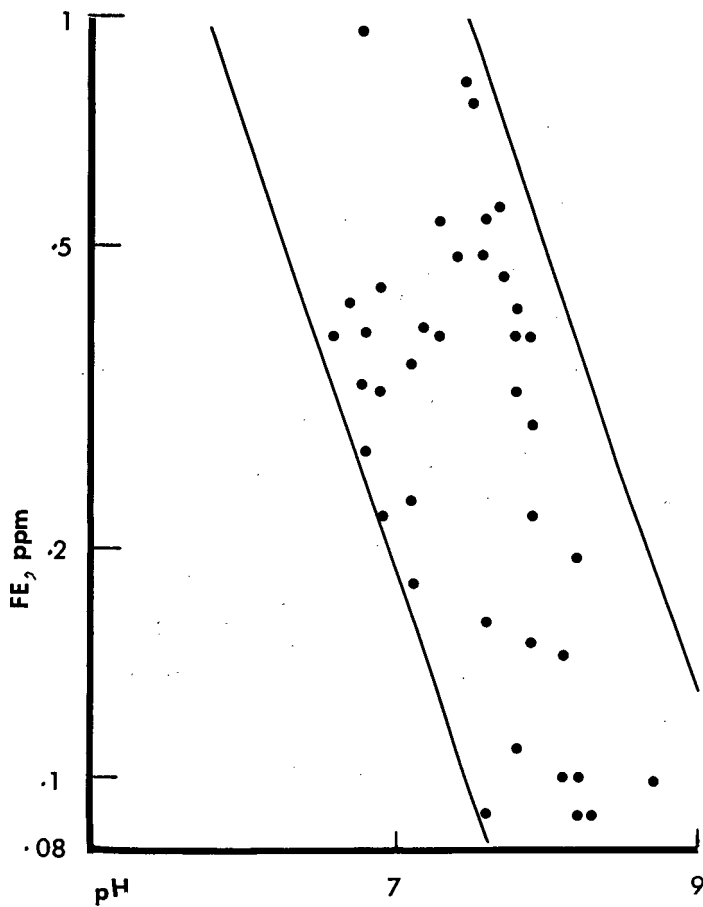


FIG. 6: Variations in total dissolved Fe with pH.

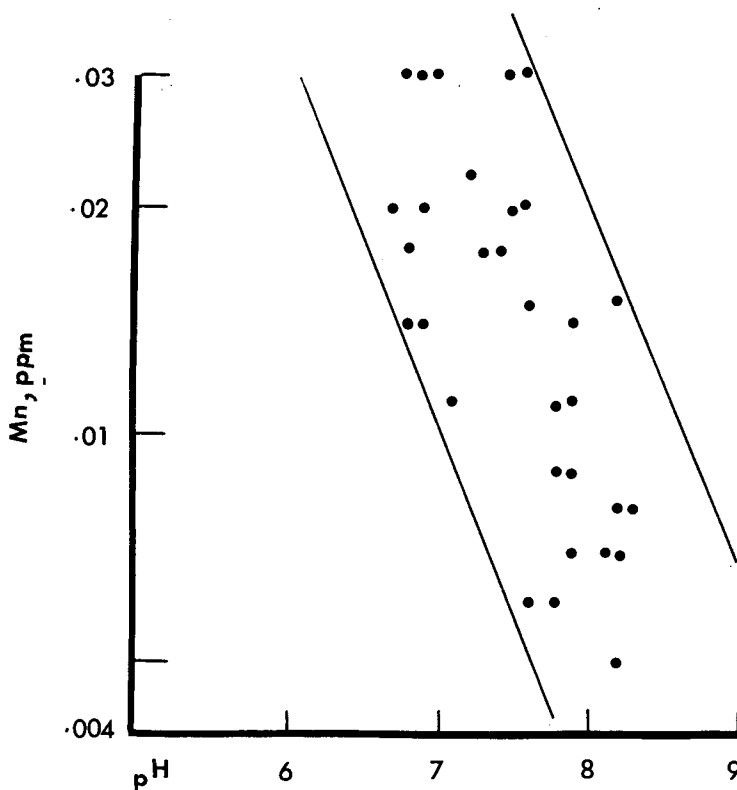


FIG. 7: Variations in total dissolved Mn with pH.

hence, no silicate mineral is likely to precipitate out of the river water.

Minor Constituents:

Fig. 6 shows the relationship between Fe content and pH of the river water. Even though a definite relationship is lacking, the range of values obtained fall within a restricted area of the plot, as shown by the lines in the figure and from this, a negative correlation can be considered to exist between pH and total dissolved Fe in St. Lawrence River water. Removal of Fe from fresh waters under oxidizing conditions is controlled by equilibrium between the water and the oxyhydroxides of Fe (Langmuir and Whitmore, 1971) and, since oxyhydroxides are only sparingly soluble, dissolved Fe in natural waters should be negligible. Organic complexes of Fe can stabilize Fe species in solutions beyond the levels dictated by the solubility products of Fe-Oxyhydroxides, Sulphides and Carbonates (Berner, 1971; Theis and Singer, 1974). In addition to the absence of any correlation between Fe and dissolved organic carbon in these river waters, and lack of any other direct evidence for the presence of organo-metallic complexes, the role of organics in the aqueous chemistry of Fe is only speculative at this stage (Plumb Jr. and Fred Lee, 1972).

In the pH range observed for St. Lawrence River water, the maximum Fe that could be in solution is less than 0.1 ppm (Hem, 1972). The observed higher values of Fe may (a) reflect analysis of fine colloidal particles of oxyhydroxides of Fe which are ubiquitous in natural aqueous systems, or (b) reflect, indirectly, the stabilizing effect of both inorganic and organic ligands in the water (Morgan and Stumm, 1965), or (c) show the effect of the total dissolved salts (salinity) (Bewer et. al. 1974).

In Fig. 7 is shown the relationship between pH and total dissolved Mn. The behaviour of Mn is very similar to that of Fe. A scatter of values at any given pH is more prominent for Mn than for Fe. The highest Mn values reported in the analyses fall within the range expected for equilibrium between hydrous Mn oxides and water (Bricker, 1965). Under identical Eh and pH conditions, Mn is stable relative to Fe in natural waters (Kruskopf, 1957).

Because of the reported intimate association of Fe and Mn in the weathering cycle (Krauskopf, 1957; Jenne, 1967), a plot of dissolved Fe versus dissolved Mn was made and is shown in Fig. 8. The observed positive correlation suggest that both Fe and Mn in St. Lawrence River water are controlled by the same factors. Because of the greater ease of formation of Fe-Organics relative to Mn-Organics (Rashid, 1972; Stumm and Morgan, 1970) and also for lack of any correlation between dissolved organic Carbon and Fe as well as Mn (plots not shown), it is unlikely that organic chelation could be a single controlling factor for both Fe and Mn in the water of the St. Lawrence River. Besides Eh and pH, source-area lithology is also important in the aqueous chemistry of Fe and Mn. The St. Lawrence River water probably

derives both Fe and Mn from the same minerals in the source area. Possible mineralogical contributions are discussed later.

Even though the positive correlation between Fe and Mn holds good at all seasons, the plot can conveniently be divided into two seasons as follows: regions of higher Fe and Mn values representing data from the month of October, and regions of lower Fe and Mn values representing data from the month of May. An insignificant number of data are available for other months and are not considered here. Colloidal Fe sorbed onto Mn oxides may possibly represent high Fe Mn values (Martin et al, 1973; Gibbs, 1972; Jenne, 1967).

Due to the non-availability of significant number of analyses, no controls could be evaluated for other trace metals.

Dissolved Organic C, N and P:

Figure 9 shows the plot of dissolved organic carbon versus total N. A negative correlation is evident and suggests that they are derived from two different sources. Dissolved organic carbon (DOC) is the result of interaction between water and the ecosystem. DOC in world average river water is 10 ppm (Livingstone, 1960) and represents nearly a tenth of the total dissolved load carried by rivers. Amino acids, proteins, lipids and carbohydrates constitute the major dissolved organic matter in natural waters. Based on thermodynamics of equilibrium systems, Thorstenson (1969) concluded that organics of higher molecular weights are unstable in most natural waters; in spite of this, higher molecular weight organics have been reported from several natural environments, - up to 5 ppm of humic matter have been reported for river waters (Martin et al, 1971) and dissolved amino acids in McKenzie river has been reported to be in the range of 15-220 ppb (Peake et al, 1972). Since the St. Lawrence River water has not been fractionated to isolate individual organic components, no specific source could be assigned for the DOC.

The source of dissolved N are NH_4^+ , amino acids, dissolved N_2 , and NO_3^- . The negative correlation between DOC and N suggests that dissolved organics containing amino functional groups are not important in the present system. Alternately, contributions from inorganic N could be in excess of that of amino acids. If the total dissolved organics are considered to be represented by DOC, then in the St. Lawrence River they are in the range expected for any fresh water bodies. Absence or insignificance or higher molecular weight organics may also explain the lack of correlation between Fe and DOC. None of the trace metals such as Zn showed any correlation with DOC. Since metal-organic complexes are formed easily with higher organic compounds such as humic acids, it appears that metals and also N chemistry in the St. Lawrence River are primarily influenced by inorganic rather than organic processes. Though the presence of dissolved Fe-P complexes are not known so far, Upchurch (1974) has shown that Fe-P inorganic complexes could exist in suspension. The positive relationship between dissolved Fe and P, shown in Figure 10, may again reflect the same source-area effects. Except over a narrow range of P-values, P content is independent of Fe content. The P-values are high compared with

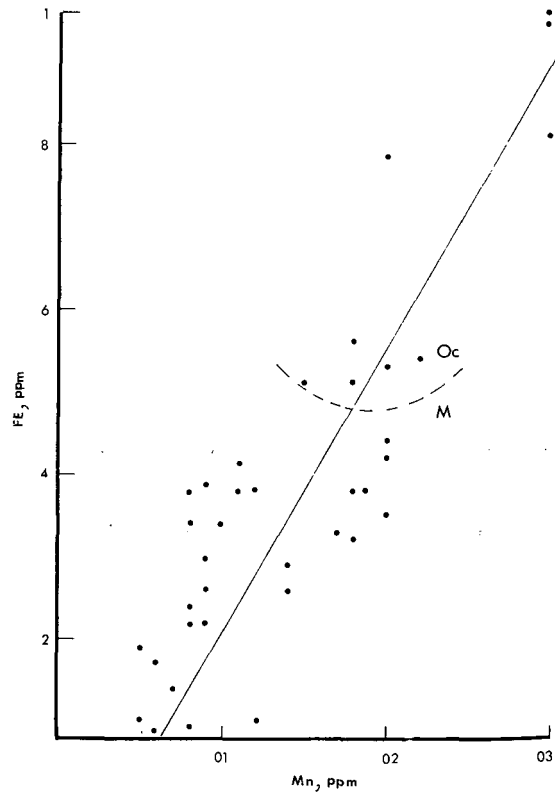


FIG. 8: Relationship between total dissolved Fe and Mn. Oc-October; M-May.

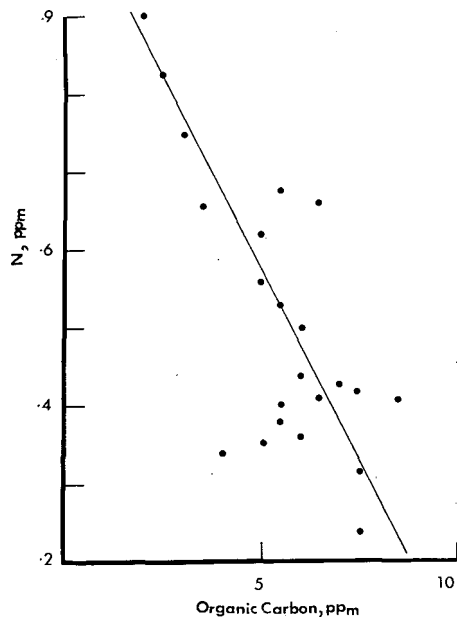


FIG. 9: Variations in total dissolved N with dissolved organic C.

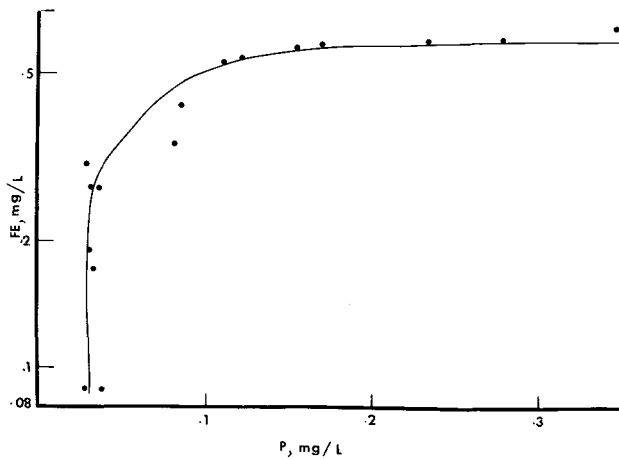


FIG. 10: Variations in P with total dissolved Fe.

those reported for estuaries (UpChurch et al, 1974) but the higher productivity in the estuary region should remove P from solution whereas in the fresh water regions, consumption by the organisms would be the limiting factor for P. St. Lawrence River water appears to be consistent with these expectations.

Water composition in the system
 $K_2O-Na_2O-MgO-CaO-Al_2O_3-SiO_2-H_2O$:

The chemistry of river water can be represented in a system diagram consisting of the above mentioned components. In the past, such diagrams have been reported (Drever, 1971; Garrels and MacKenzie, 1971; Jacks, 1973; Helgeson and MacKenzie, 1971; Kramer, 1967, 1968; Norton, 1974). Hydrolysis reactions of silicates or silicate-silicate reactions can be written as a function of three variables - a cation, pH and dissolved silica. Equilibrium reactions between any two silicate minerals can thus be represented in an orthogonal plot consisting of two variables - Cation/Hydrogen, and SiO_2 .

In the diagrams shown in Fig. 11 the composition of St. Lawrence River water has been plotted in stability diagrams of silicate minerals expressed as a function of the three above-mentioned variables. Except for a single analysis, the river water falls in the gibbsite and chlorite area. Kramer (1967, 1968) has reported that Lake Ontario samples, from winter and summer months, fall in the gibbsite and kaolinite areas on similar plots. Ruthford (1972) has observed illite, chlorite and mixed-layer clays in the bottom sediments of the St. Lawrence River near Kingston, Ontario at the point where the river flows from the Great Lakes. Gibbsite has not been reported anywhere from the St. Lawrence basin sediments. Even in the St. Lawrence estuary and the Gulf of St. Lawrence, the mineralogy of sediments reported so far consist of chlorite, illite and mixed-layer clays with occasional kaolinite but no gibbsite (d'Anglejan and Smith, 1973; Beltagy, 1974). Thus the observed mineralogy is different from that predicted for the river based on water chemistry. Possible causes are discussed later.

Assuming (1) the average composition of the river water is that at Lewis, Quebec, (2) hydrolysis data for silicate minerals given by Helgeson (1969), Kramer (1968) and (3) following the methods of Helgeson and McKenzie (1971), the silicate system can be treated simply as a two component system consisting of pH and silica. Fig. 12 shows such a plot. The water compositions are plotted in the stability diagram as before, and the predicted mineralogy is the same now as was the case when the system was plotted in a seven-component system (Fig. 11).

DISCUSSION

Hydrological data for the St. Lawrence basin are summarized in Table 1. Nearly 60 percent of the daily average drainage at Quebec City (at the end of the fresh water region) can be accounted

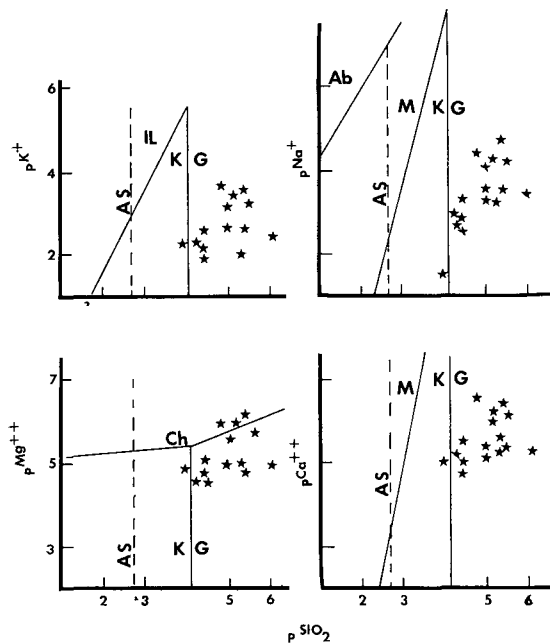


FIG. 11: Water composition in the multicomponent system.
 $pK^+ = -\text{Log}(K^+/H^+)$; $pNa^+ = -\text{Log}(Na^+/H^+)$;
 $pMg^{++} = -\text{Log}(Mg^{++}/H^+)$; $pCa^{++} = -\text{Log}(Ca^{++}/H^+)$; $pSiO_2 = -\text{Log}(H_4SiO_4)$;
 Ab-Albite; AS-Amorphous Silica saturation; M-Montmorillonite; IL-Illite; Ch-Chlorite; K-Kaolinite; G-Gibbsite.

Table 1. Data on St. Lawrence River

Source	Drainage Km ²	Station	Discharge m ³ /sec	TDS ppm
<u>1. Sedimentary Basin</u>				
Lake Ontario (a)	63973	--	6624	194
Richelieu (b)	19300	Chambly	318	87
St. Francois (b)	8300	Drummondville	2.3	--
<u>2. Canadian Shield Basin</u>				
St. Maurice (c)	42000	Centre de Grand Mere	708	20
Ottawa (d)	89600	Chat Falls	703	--
St. Lawrence (e)	766000	Quebec City	10350	160

- a. Data from Ragtozkie, 1974.
 - b. Data from Annuaire Hydrologique, 1970.
 - c. Data from Can. Comm. Int. Hy. Decade, 1972. TDS data from Saguenay River. (Loring and Nota, 1973).
 - d. Data from Can. Comm. Int. Hy. Decade, 1972.
 - e. Data from Padre, 1948 and quoted in Loring and Nota (1973).
- TDS - Total Dissolved Solids.

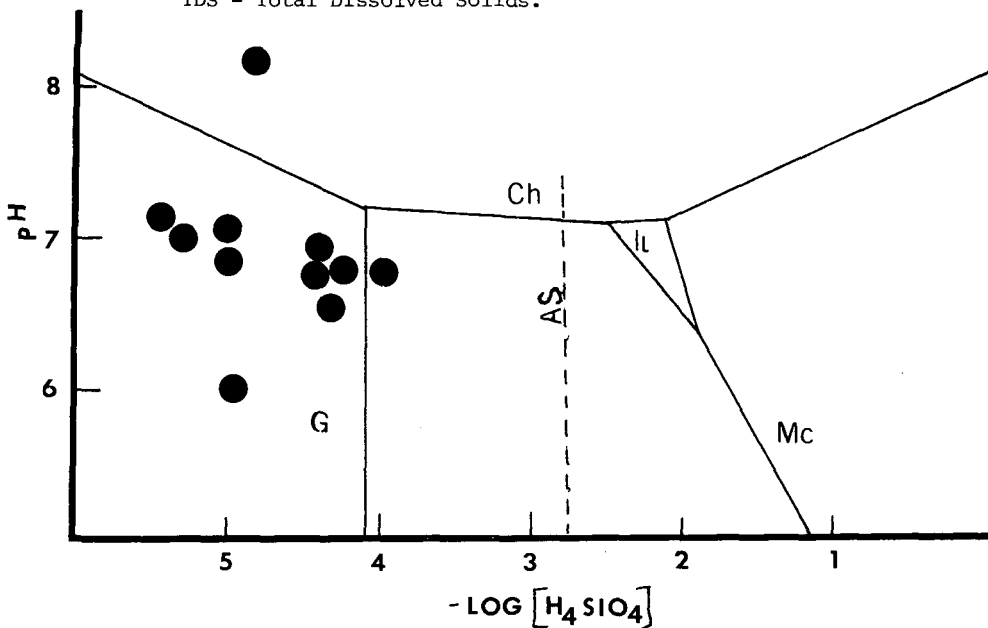


FIG. 12: Water composition plotted on an orthogonal diagram. Abbreviations same as above. Mc-Microcline.

for by water draining a sedimentary basin. This sedimentary water carries a high percentage of dissolved load compared with waters from the Canadian Shield area. Because of the large contribution of Lake Ontario to the drainage of the St. Lawrence river and also because of the presence of the Niagara dolomites in regions adjoining Lake Ontario, St. Lawrence water should be a carbonate water, based on the classification of Garrels and MacKenzie (1971). The high percentage of HCO_3^- in the river water is consistent with the drainage lithology. Only a small increase in the total dissolved load of Lake Ontario in recent years has been reported (Ragotzkie, 1974). Because of the close similarity in the water composition of Lake Ontario and of the river near Quebec City, it can be stated that the

river-water chemistry is essentially the same all along the fresh-water region. Of the four main tributaries, discharge from the hard-rock terrain (Ottawa River and St. Maurice River) is nearly 4.5 times that from the soft-rock basin (Richelieu River and St. Francois River). Thus the water with a high dissolved load from Lake Ontario is diluted by the smaller contribution from the hard-rock regions before the river reaches Quebec City. Thus, dilution, along with possible entrapment of the load at Lake St. Peter half way between Montreal and Quebec City, with possible inorganic precipitations probably account for a small net loss of dissolved load at the end of the fresh-water region of the river (See Table 1).

Except for the Na^+ , SO_4^{--} and Cl^{--} contents, the individual elemental composition of the water is not different along the course of the river. While the Na^+ and Cl^{--} contents can be modified either by additions through winter salts or Champlain Sea sediments, changes in SO_4^{--} content possibly reflect industrial additions. Even at the highest SO_4^{--} values, the St. Lawrence river is under-saturated with respect to all metal-Sulphate phases and hence the SO_4^{--} will remain in the water column of the river. Rutherford (1972) reported increased concentration of SO_4^{--} in the interstitial waters of bottom sediments near Kingston but such an increase may be entirely due to diagenesis.

Major ionic ratios suggest that their concentrations are primarily controlled by weathering processes, based on Gibbs' (1970) classification of river compositions. Precipitation does not appear to have significant effects on the water chemistry of the St. Lawrence River. Because of the dominance of the sedimentary basin, mineral-water reaction, rather than physical weathering is likely to contribute to the water chemistry.

Since the reported composition of suspended matter in the estuary and the gulf predominantly consist of detrital silicates, the chlorite and gibbsite expected to form under equilibrium conditions cannot be distinguished from the detrital phases even if they form through water-mineral reactions. However, gibbsite has never been reported and this may be due to its presence in amorphous forms such as bohemite or both gibbsite and kaolinite react with water to become illite. Such a halmyrolitic reaction is unlikely because of the detrital nature of the suspensoids. The author is currently experimenting with gibbsite and bohemite reactions with fresh water to monitor changes that may be relevant to the St. Lawrence River water chemistry. The author is also studying suspended matter from the river along a transect from the fresh water to completely salt-water region to identify any halmyrolitic reaction taking place in the river. The results should be available in the near future.

The pH control of the chemistry of Fe and Mn in natural water is well known (Hem, 1972), and data from the river indicate that Fe and Mn behave in similar fashion and that both are derived from the same source mineral. Rutherford (1972) mentioned the presence of Mn-rich regions in the Great Lakes and the same could also be the source for Fe. Cronan and Thomas (1972) located ferromanganese concretions in the Great Lakes and solution of these nodules, under favourable conditions, may provide both Fe and Mn to the St. Lawrence River water. Fe values are locally high, often 0.9 ppm; such high values reflect possible analyses of colloidal Fe-hydroxides. The river water is mostly saturated with respect to $\text{Fe}(\text{OH})_3$ and the high values may also reflect metastable equilibrium.

In lieu of the large sedimentary drainage basin, the negative correlation between C and N is surprising. Organic matter in sedimentary rocks is common and river-water interaction should mobilize some of them. The Great Lakes have a

variety of dissolved organics including those with higher molecular weights (Ragotzkie 1974) and their near paucity in the river suggest some removal processes before the river drains out of the Lakes.

CONCLUSION

In spite of the small drainage of the St. Lawrence River compared with other major rivers of the world, the total dissolved load carried by the river into the North Atlantic may be significant in mass transfer and budget studies. The major ionic composition of the river is essentially the same throughout the fresh-water region. A small net loss if dissolved load is observed which may be due to the presence of Lake St. Peter.

The water chemistry of the river is primarily controlled by rock domination in the water shed. The sedimentary part of the drainage basin contributes the major ions which are slightly diluted by large-volume discharge from the Canadian Shield area. The predicted mineral equilibria from the water chemistry has not yet been observed from the actual mineralogical studies of the suspended matter.

REFERENCES

- D'ANGLEJAN, B.F., and SMITH, E., 1973, Distribution, transport and composition of suspended matter in the St. Lawrence Estuary, *Can. Jour. Ear. Sc.*, 10, p. 1380-1396.
- ANNUAIRE HYDROLOGIQUE, 1970, Ministry de Rechesse Naturelles, Govt. du Quebec, 409 pp.
- BELTAGY, A., 1974, Geochemistry of some recent marine sediments from the Gulf of St. Lawrence, Ph.D. Thesis, McGill University, Montreal, Quebec, Canada, 203 pp.
- BERNER, R.A., 1971, *Chemical Sedimentology*, McGraw Hill Pub, N.Y. 240 pp.
- BEWER, J.M., MCCAULAY, I.A., and SUNDBY, B., 1974, Trace Metals in the Gulf of St. Lawrence, *Can. Jour. Ear. Sc.*, 11, p 939-950.
- BRICKER, O.P., 1965, The System $\text{Mn}-\text{O}_2-\text{H}_2\text{O}$ at 25°C and one atmospheric pressure, *Am. Min.* 50, p 1296-1356.
- CANADIAN COMMISSION FOR I.H.D., 1972, Discharge of selected Canadian rivers, Information Canada, Ottawa, 338 pp.
- CRONAN, D.S., and THOMAS, R.L., 1972, Geochemistry of ferromanganese concretion in Lake Ontario. *Bull. Geol. Soc. Am.*, 83, p 1493-1502.
- DREVER, J.L., 1971, Chemical weathering in a sub-tropical igneous terrain. *Jour. Sed. Petrology*, 41, p 951-961.
- EDWARDS, A.M.C., 1973, The variations of dissolved constituents with discharge in some Norfolk rivers, *Jour. Hydrology*, 18, p 219-242.

- GARRELS, R.M., and MACKENZIE, F.T., 1971, The evolution of sedimentary rocks, Norton Pub, N.Y.
- _____, and CHRIST, C.L., 1965, Solution, minerals and equilibria. Harper and Row, Pub, N.Y.
- GIBBS, R.J., 1967, Geochemistry of the Amazon River, Bull. Geol. Soc. Am. 78, p 1203-1232.
- _____, 1970, Mechanisms controlling world water chemistry, Science, 170, p 1089-1090.
- _____, 1972, Water chemistry of the Amazon River, Geochem. Cos. Acta., 36, p 785-789.
- GOVE, G.W., and GELLMAN, I., 1972, Industrial wastes: paper and allied products, Jour. W.P.C. Fed., 44, p 1046-1048.
- GOVERNMENT OF INDIA, 1955, A study of quality of water in important rivers in India, C.W.P.R. Station. 22, p 111-116.
- GROVE, A.T., 1972, The dissolved and solid load carried by some west African river, Jour. Hydrology, 16, p 277-300.
- HANDA, B.K., 1972, Geochemistry of the Ganga River Water, Indian Geohydrology, 8, p 71-78.
- HELGESON, H.T., 1969, Thermodynamics of hydrothermal systems at elevated T and P. Am. Jour. Sc. 267, p 729-804.
- _____, and MACKENZIE, F.T., 1971, Silicate-sea water equilibrium. Deep-Sea Research, 17, p 877-892.
- HEM, J.D., 1972, Chemical factors affecting the availability of Fe, Mn. Bull. Geol. Soc. Am., 83, p 443-450.
- INMAN, C.L., and NORDSTROM, C.E., 1971, Tectonic and morphological classifications of coasts, Jour. Geol., 79, p 1-21.
- JACKS, G., 1973, Chemistry of ground water from a district in S. India, Jour. Hydrology, 18, p 185-200.
- JENNE, E.A., 1967, Controls on Fe, Mn, Co, Ni, Cu and Zn in soils and water, Am. Chem. Soc. Spec. Pub., 73, p 337-387.
- KRAMER, J.R., 1967, Equilibrium models and composition of the Great Lakes, Am. Chem. Soc. Spec. Pub. 67, p 243-251.
- _____, 1968, Mineral water Equilibria in silicate weathering, Proc. 13th Int. Geol. Cong., 10, p 149-160.
- KRUSKOPF, K., 1957, Factors affecting the concentration of 13 rare elements, Geochem, cos. Acta., 12, p 61-92.
- LAGMUIR, D., and WHITMORE, D., 1971, Variations in the stability of precipitated ferric, oxyhydroxides, Am. Chem. Spec. Pub. 106, p 209-234.
- LIVINGSTONE, D.A., 1960, Chemical composition of rivers and lakes, U.S.G.S. Prof. Pub., 440G.
- LORING, D.H., and NOTA, D.J.G., 1973, Morphology and sediments of Gulf of St. Lawrence, Bull. Fish. R.B. Canada, 182, 147 pp.
- MCKEAGUE, J.A., and CLINE, M.G., 1963, Silica in soil solutions. Can. Jour. Soil Sc., 43, p 70-82.
- MACKENZIE, F.T., and GARRELS, R.M., 1965, Silicates: reactivity with sea water, Science, 150, p 157.
- MARTIN, D.F., DOIG, M.T., and PIERCE, R.H., 1971, Naturally occurring chelators, Fl. Dept. N.R. Prof. Paper, 12, p 1-52.
- MARTIN, J.M., KULBICKI, G., and DEGROOT, A.J., 1973, Terrigenous supply of radioactive and stable elements, Hydrogeochemistry, 1, p 463-484.
- MORGAN, J., and STUMM, W., 1965, Role of multivalent oxides in limnological transformations, Adv. W.P. Res., 2, p 103-131.
- MYTELKA, A.I., 1972, Industrial wastes: chemical and allied products, J.W.P. Con. Fed., 44, p 1113-1114.
- NORTON, D., 1974, Chemical mass transfer in Rio Tanama system Geochem. Cos. Acta., 38, p 267-277.
- PEAKE, E., BAKER, B.L., and HODGSON, G., 1972, Hydrogeochemistry of the surface waters of Mackenzie River basin, Geochem. Cos. Acta, 36, p 867-83.
- PLUMB, R.H., and FRED, LEE, G., 1973, A Note on Fe-organic relationship in natural water. Water Res., 7, p 581-585.
- RAGOTZKIE, R.A., 1974, The Great Lakes rediscovered, Am. Scientist, 62, p 454-464.
- RASHID, M.A., 1972, Role of quinone groups in solubilizing and complexing metals, Chem. Geol. 9, p 241-248.
- RUTHFORD, G.K., 1972, The clay mineralogy and some properties of bottom sediments of the St. Lawrence River. Can. Jour. Earth Sc., 8, p. 1670-1676.
- SIEVER, R., 1962, Silica solubility, 0-200 C and the diagenesis of siliceous sediments. Jour. Geol., 70, p 127-150.
- _____, and WOODWARD, N., 1973, Sorption of silica by clay minerals, Geochem. Cos. Acta., 37, p 1851-1880.
- STUMM, W., and MORGAN, J., 1970, Aquatic Chemistry, Wiley Interscience Pub, N.Y. 583 pp.
- THEIS, T.W., and SINGER, P.C., 1974, Complexing of Fe (II) by organic matter and its effects on Fe (II) oxygenation, Env. Sc. Tech., 8, p 569-573.
- UPCHURCH, J.E., EDZWALD, J.K., and O'MELIA, C.R., 1974, Phosphates in sediments of Pamlico Estuary, Env. Sc. Tech. 8, p 56-58.